

Fig. 2-3. $p-v-T$ surface for an ideal gas.

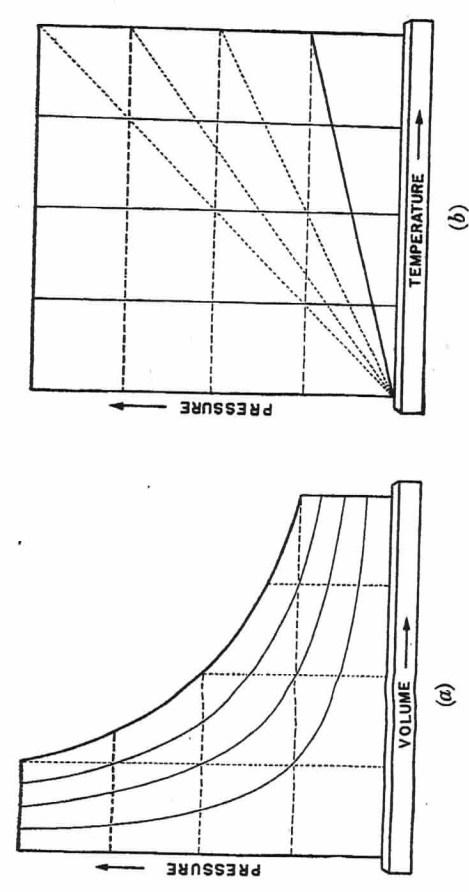


Fig. 2-4. Projections of ideal gas $p-v-T$ surface on (a) the $p-v$ plane, (b) the $p-T$ plane.

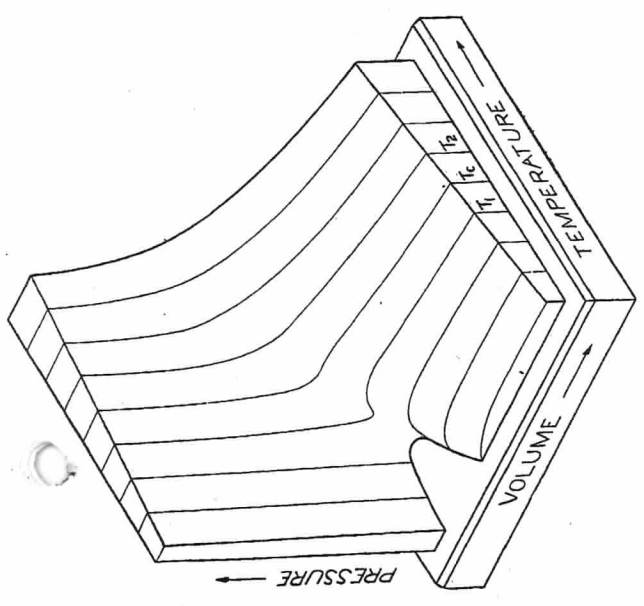


Fig. 2-5. $p-v-T$ surface for a van der Waals gas.

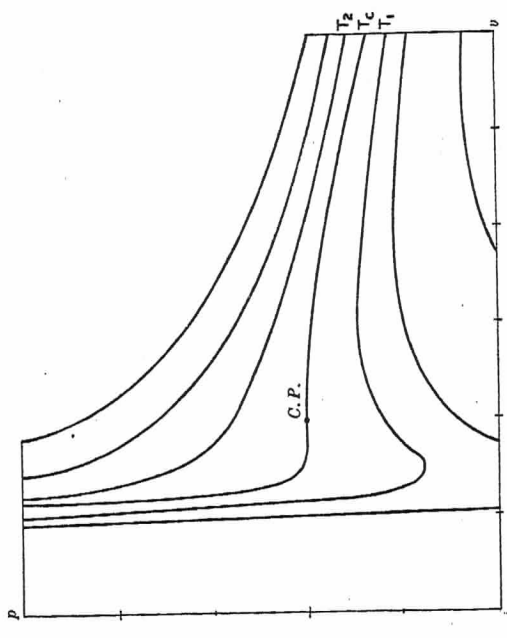


Fig. 2-6. Isotherms of a van der Waals gas.

$$pv^3 - (pb + RT)v^2 + av - ab = 0$$

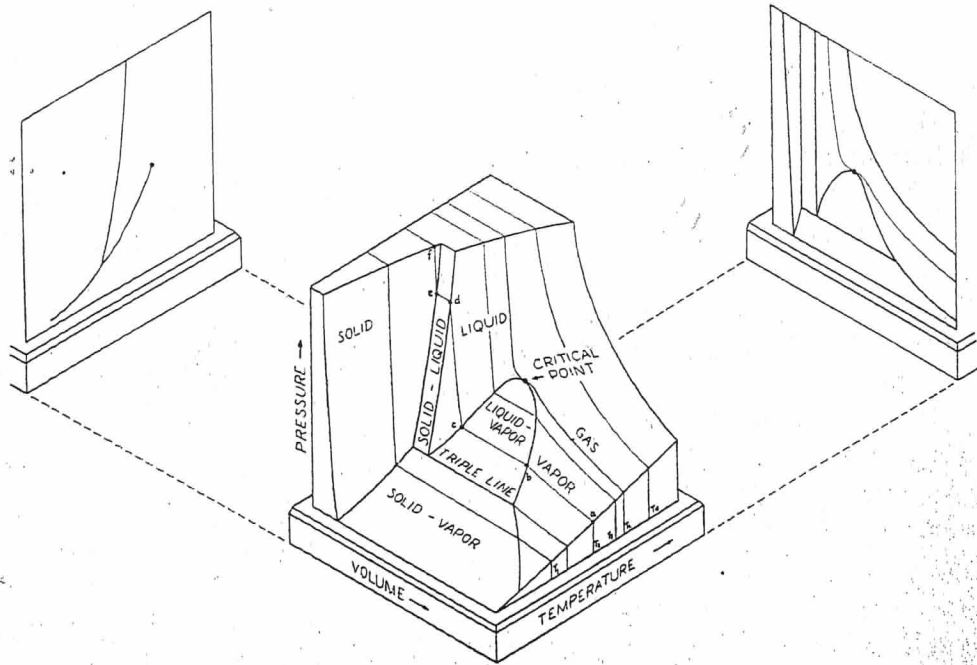


FIG. 6-5. Projection of $p-v-T$ surface on the $p-T$ and $p-v$ planes.

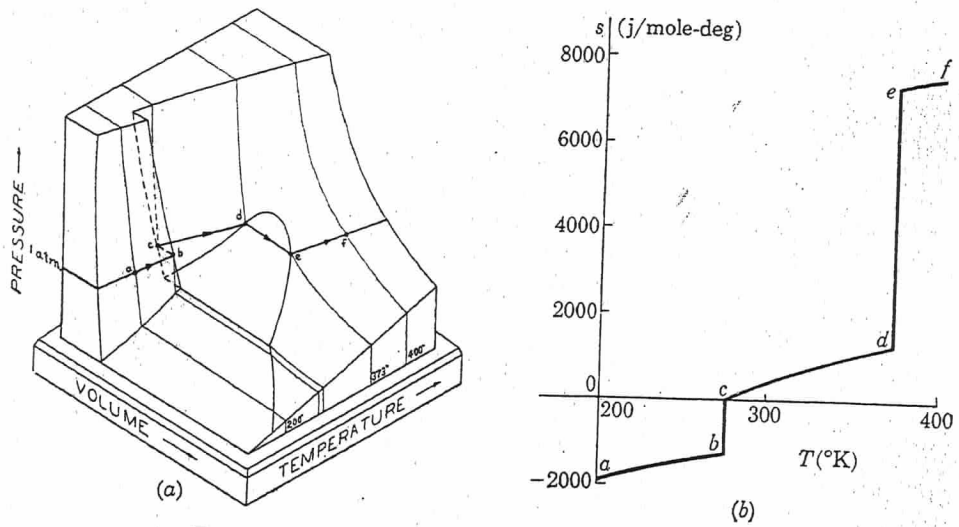


FIG. 8-3. Entropy changes in an isobaric process.