

FIG. 2-3. p - v - T surface for an ideal gas.

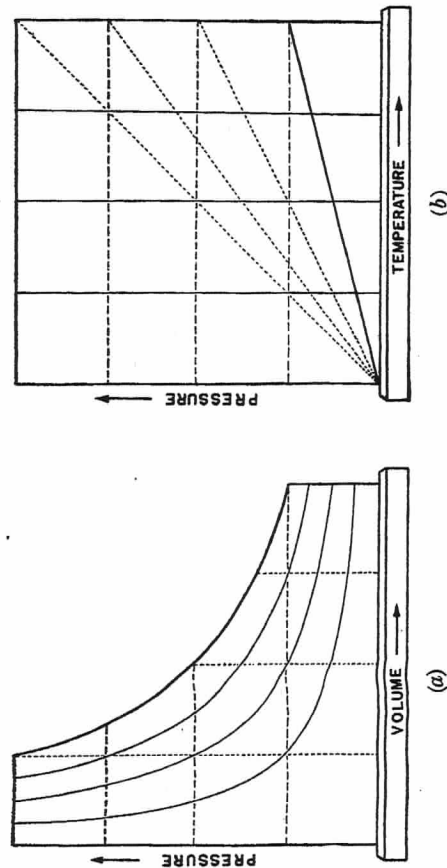


FIG. 2-4. Projections of ideal gas p - v - T surface on (a) the p - v plane, (b) the p - T plane.

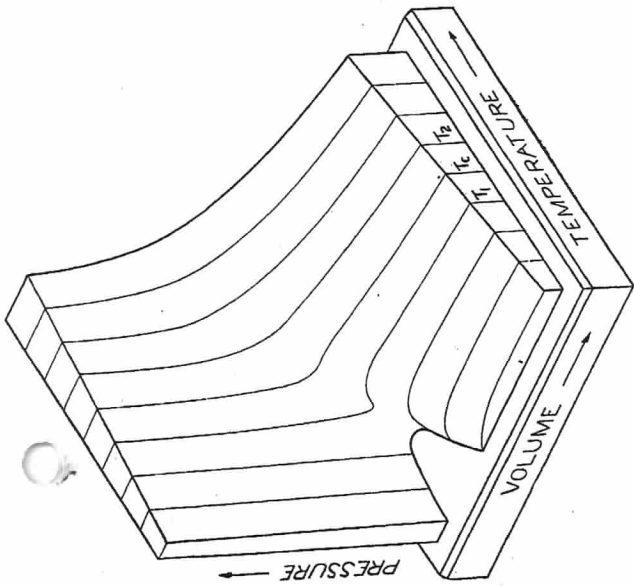


FIG. 2-5. p - v - T surface for a van der Waals gas.

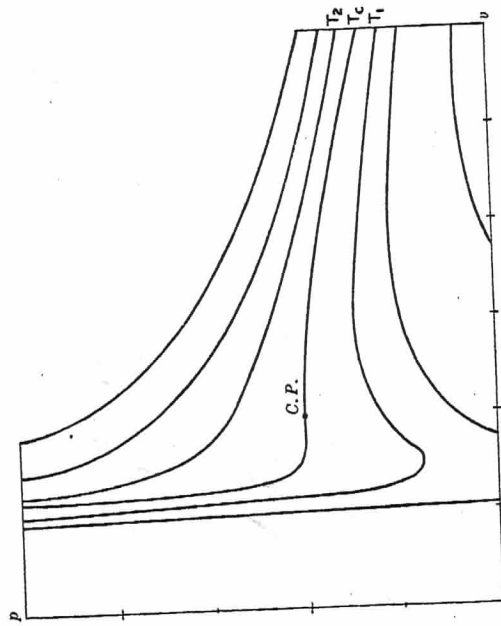


FIG. 2-6. Isotherms of a van der Waals gas.

$$pv^3 - (pb + RT)v^2 + av - ab = 0$$

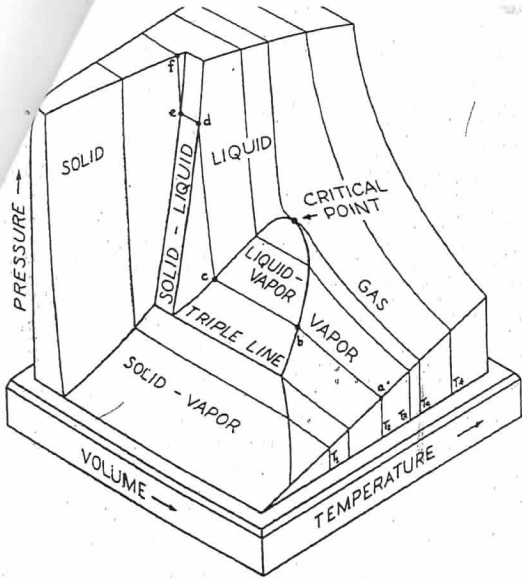
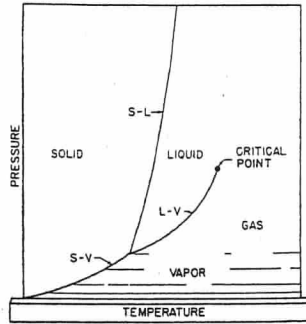
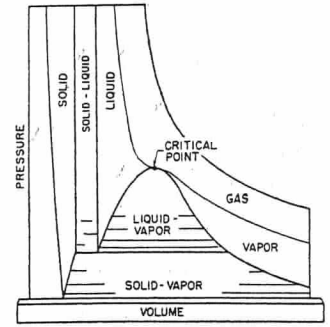


Fig. 6-1. $p-v-T$ surface for a substance that contracts on freezing.

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(a)



(b)

Fig. 6-3. Projections of the surface in Fig. 6-1 on the $p-T$ and $p-v$ planes.

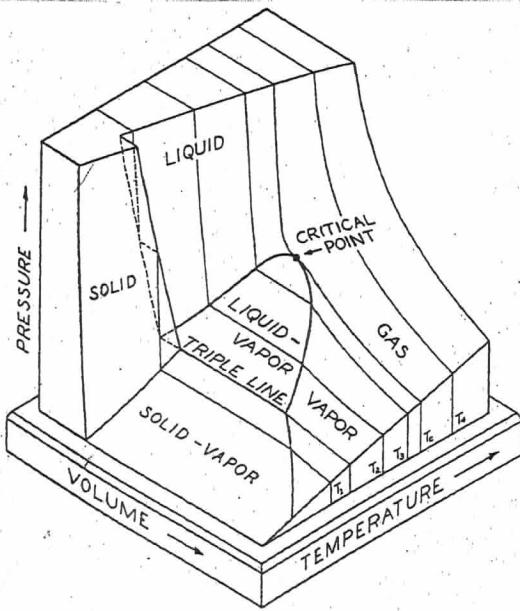
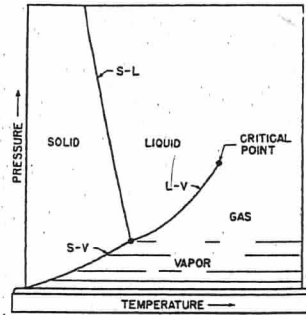
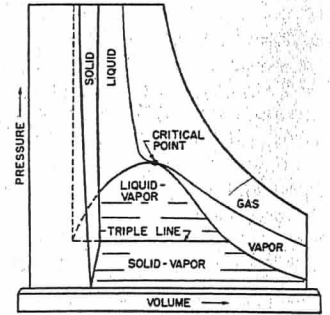


Fig. 6-2. $p-v-T$ surface for a substance that expands on freezing.



(a)



(b)

Fig. 6-4. Projections of the surface in Fig. 6-2 on the $p-T$ and $p-v$ planes.

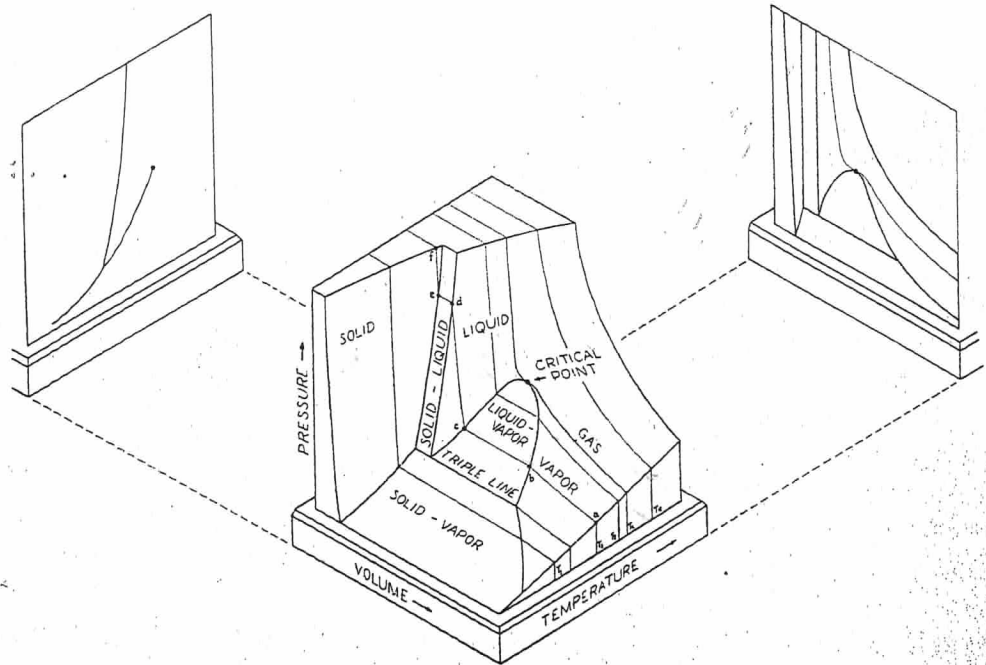


FIG. 6-5. Projection of p - v - T surface on the p - T and p - v planes.

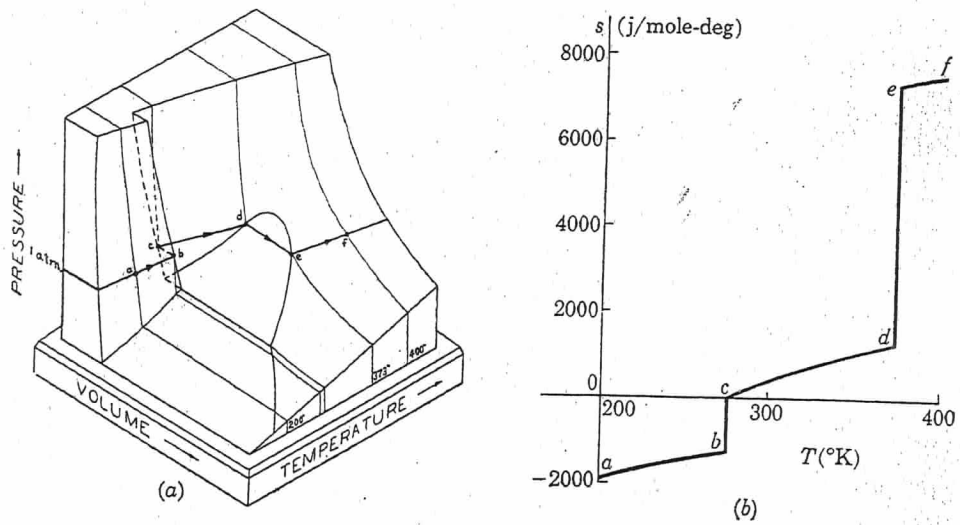


FIG. 8-3. Entropy changes in an isobaric process.